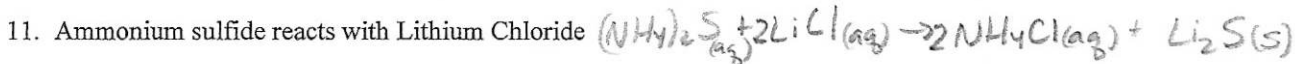
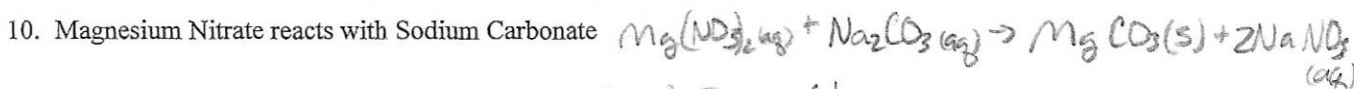
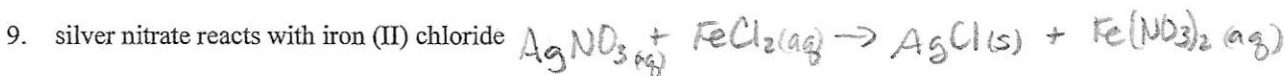
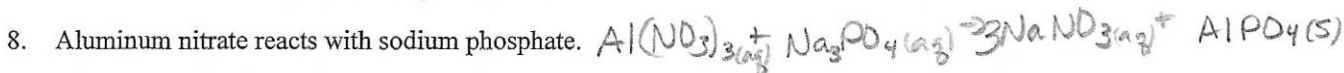
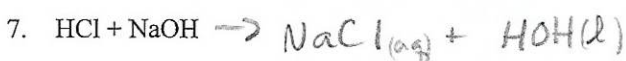
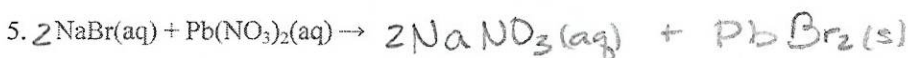
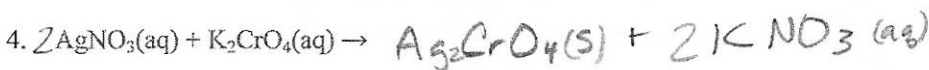
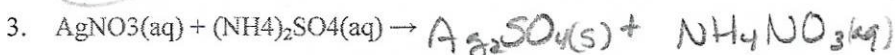
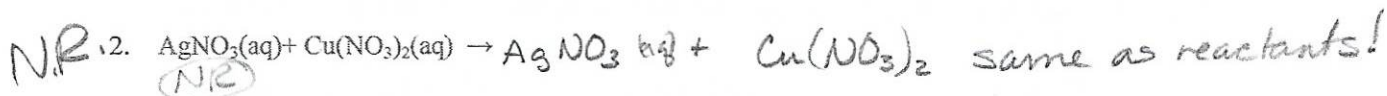
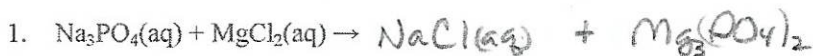


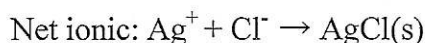
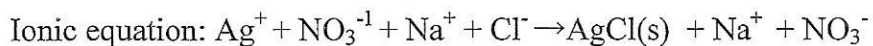
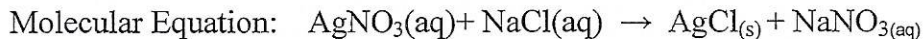
Name
Chemistry
Solubility Reactions #1

Molecular Equation: $AB + CD \rightarrow CB_{(aq)} + AD_{(s)}$	
Insoluble = precipitate Soluble = aqueous	Quick Check Na^+X K^+X NH_4^+X XNO_3^-
***Need at least 1 insoluble product or water	Always Soluble

- Determine potential products
- Indicate if the product will be soluble(s), insoluble (aq), or Liquid (H₂O(l))
- If there are no "NEW" products indicate no reaction.

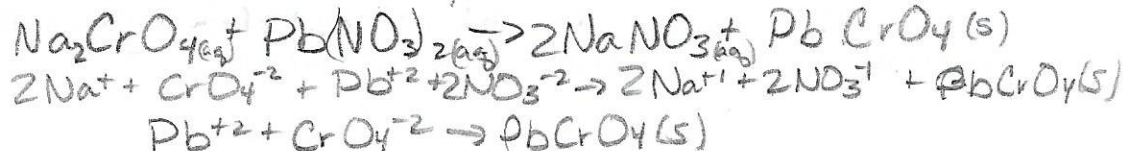


Name
Chemistry
Solubility Reactions #2

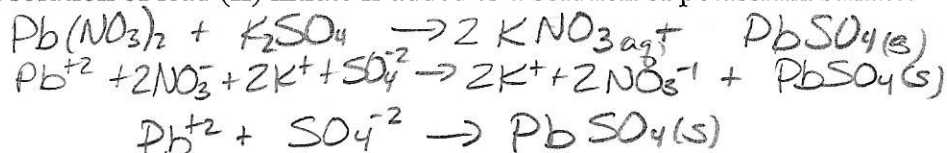


Complete Molecular, ionic, and net-ionic equations for the following

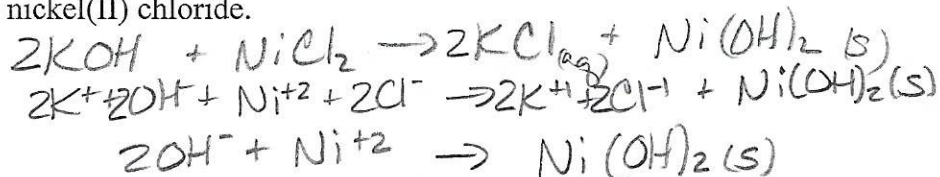
1. Solutions of sodium chromate and lead nitrate are mixed.



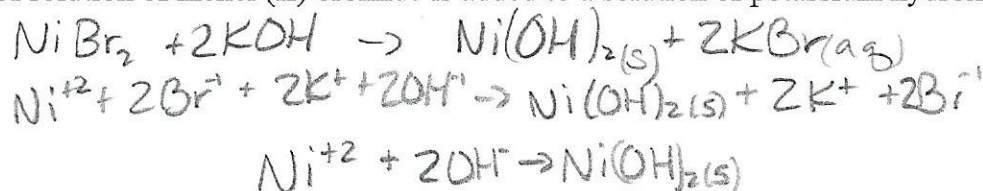
2. A solution of lead (II) nitrate is added to a solution of potassium sulfate.



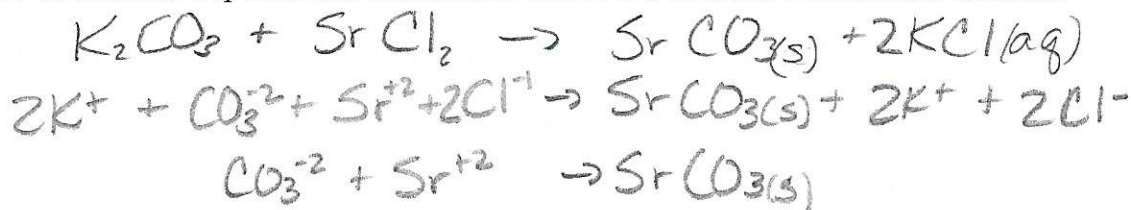
3. Excess concentrated potassium hydroxide solution is added to a solution of nickel(II) chloride.



4. A solution of nickel (II) bromide is added to a solution of potassium hydroxide.



5. A solution of potassium carbonate is added to a solution of strontium chloride.



6. A solution of copper (II) chloride is added to a solution of sodium sulfide.

